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Volume (V) Temperature (T) C. Find: The pressure of the steam.

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Chapter 3: Pure Substances Phase Change and Property Diagrams  
In this chapter we consider the property values and relationships of  
a pure substance (such as water) which can exist in three phases –  
solid, liquid and gas.

## Chapter 3: Pure Substances – Thermodynamics

The change in internal energy can be found from the first law of  
thermodynamics:  $U = Q - W = (3.5 \times 10^5 \text{ J}) - (2.1 \times 10^5 \text{ J}) = 0.9$   
 $\times 10^5 \text{ J} = 90 \text{ kJ.}$  A gas in a cylinder is kept at a constant pressure  
of  $3.5 \times 10^5 \text{ Pa}$  while 300 kJ of heat are added to it, causing the gas  
to expand from  $0.9 \text{ m}^3$  to  $1.5 \text{ m}^3$ .

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3-5 3-23 Problem 3-22 is reconsidered. The missing properties of  
water are to be determined using EES, and the solution is to be  
repeated for refrigerant-134a, refrigerant-22, and ammonia.

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Chapter 3: THERMODYNAMICS. -Thermodynamics is the study of the relationship between the energy transformation in the system and other physical quantities such as temperature, pressure and volume (P, V, T). -A thermodynamic equation of state is a mathematical relationship of the thermodynamic or state variables, such as pressure, volume and temperature.

## Chapter 3: THERMODYNAMICS

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solution manual Thermodynamics: An Engineering Approach ...  
Chapter 3-3 Heat transfer is energy in transition due to a temperature difference. The three modes of heat transfer are conduction, convection, and radiation. Conduction through Plane Walls Conduction heat transfer is a progressive exchange of energy between the molecules of a substance. Fourier's law of heat conduction is  $Q = -kA \frac{dT}{dx}$  here Q!

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