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ratios are used

to solve

problems such as

how many moles

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dioxide CO_2
would be
produced from 6
25 moles of
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moles 0.23 mol
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$4\text{NH}_3(\text{g}) + 6\text{NO}$
 $(\text{g}) \rightarrow 5\text{N}_2(\text{g}) +$
 $6\text{H}_2\text{O}(\text{g})$ How
many moles of
each reactant

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were there if
13.7 moles of
 $\text{N}_2(\text{g})$ is
produced? $\times 4$
moles $\text{NH}_3(\text{g}) =$
10.96 moles NH_3
 $(\text{g}) \times 6$ moles NO
 $(\text{g}) = 16.44$
moles $\text{NO}(\text{g})$ So
we have 10.96
moles $\text{NH}_3(\text{g})$ and
16.44 moles NO
 (g) . Problem :
What is the mass

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of 2 moles of

H₂S ?

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problems. 1.
Start with your
balanced
chemical

equation. 2.

Convert the
given mass or
number of
particles of a
substance to the
number of moles.

3.

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+ O₂ + CO₂ + H₂O

= Cu₂ (OH)₂CO₃.

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Answers Step 1:

Balance The

Equation &

Calculate the

Ratios. $2\text{Al} : 6\text{HCl}$

$(1 : 3)$ $2\text{Al} : 2\text{AlCl}_3$

$3(1 : 1)$ $2\text{Al} : 3\text{H}_2$ 2

$(1 : 1.5)$ Step 2:

Find the Moles

of the Given.

0.87 moles of

aluminum are

reacted with

hydrochloric

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(Answer Key) NH

4 NO 3 N 2 O + 2

H 2 O e. 4 CH 3

NH 2 + 9 O 2 4

CO 2 + 10 H 2 O

+ 2 N 2 f.

Cr(OH) 3 + 3

HClO 4 Cr(ClO 4)

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Write the
balanced
chemical

equations of
each reaction:

a.

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industrially by
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using the mole

ratio A, given

in the balanced

chemical

equation Mass of

substance A

Amount in mol of

substance A

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substance B

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